

**ATOMIC ENERGY CENTRAL SCHOOL-  
KUDANKULAM**

**Worksheet –Module-3/4**

**Subject-Chemistry**

**Class-X**

**Lesson No.- Chapter 1- Chemical Reactions and Equations**

**Name of the topic – Displacement, Double Displacement,  
Neutralisation Reaction**

**Total Marks-10x3=30**

- (1) What happens when a piece of
  - (a) Zinc metal is added to copper sulphate solution.
  - (b) Aluminium metal is added to dilute hydrochloric acid
  - (c) Silver metal is added to copper sulphate solution
- (2) On adding a drop of barium chloride solution to an aqueous solution of sodium sulphite, white precipitate is obtained.
  - (a) Write a balanced chemical equation of the reaction involved.
  - (b) What other name can be given to this precipitation reaction.
  - (c) On adding dilute hydrochloric acid to the reaction mixture white precipitate disappears, why?
- (3) Why non metals do not displace hydrogen from dilute acids. Explain.
- (4) When silver nitrate is added to seawater, a white precipitate is formed. Explain.
- (5) Fill in the blanks.
  - (a) When an element displaces another element from its compound, a \_\_\_\_\_ reaction occurs
  - (b) Precipitation reactions produce \_\_\_\_\_ salts.
  - (c) Those reactions, in which two compounds react by an exchange of ions to form two new compounds, are called \_\_\_\_\_ reactions
- (6) A solution of  $\text{AgNO}_3$  is mixed with a solution of  $\text{K}_2\text{S}$ . Write the molecular and net ionic equations illustrating the reaction.

(7) Give equation for the neutralisation reaction for the formation of the following salts.

(a) Sodium nitrate

(b) Magnesium chloride

(c) Calcium sulphate

(8) Milk of magnesia is used as an antacid. Comment on the statement. Give equations in support of your answer.

(9) Complete the following reactions.

(a)  $\text{H}_2\text{SO}_4 + \text{NH}_4\text{OH} \rightarrow$

(b)  $\text{HCl} + \text{Ca}(\text{OH})_2 \rightarrow$

(c)  $\text{CH}_3\text{COOH} + \text{NH}_4\text{OH} \rightarrow$

(10) What is the end product of neutralisation reaction? Give one important use of neutralisation reaction in daily life.

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